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Write the word equations below as chemical equations and balance. I. zinc 0 lead (II).

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Reactions Directions: Classify the reactions below as synthesis, decomposition, single replacement (cationic or anionic) or double replacement. Page 8/29

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(Answer Key) Determine the oxidation number of the elements in each of the following compounds: a.  $\text{H}_2\text{CO}_3$  H: +1, O: -2, C: +4 b.  $\text{N}_2$  N: 0 c.  $\text{Zn}(\text{OH})_4^{2-}$  Zn: 2+, H: +1, O: -2 d.  $\text{NO}_2$  N: +3, O: -2 e.  $\text{LiH}$  Li: +1, H: -1 f.  $\text{Fe}_3\text{O}_4$  Fe:  $+8/3$ , O: -2; Identify the species being oxidized and reduced in each of the following reactions:

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Questions. Question 1. Define electrochemical cell. Answer: Electrochemical cell is a device in which the redox reaction is carried indirectly and the decrease in free energy appears as electrical energy.

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## ELECTROCHEMISTRY CROSSWORD

ACROSS 4. Unit of electrical potential . 6.  
Electrode where oxidation takes place .  
7. Both atoms and \_ must be . balanced  
in a redox equation. 9. The anode in  
an.electrochemical cell has this charge.  
10. Gain of electrons . 12. Voltage of an  
electrochemical cell when it reaches  
equilibrium 13. A substance that is

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oxidized . is . the \_ agent.

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oxidation number of the elements in  
each of the following compounds: a. H<sub>2</sub>

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Word Equations Answer Key 1. Zinc and Lead (II) nitrate react to form Zinc Nitrate and Lead. + Pb Single Replacement 2. Aluminum Bromide and Chlorine gas react to form Aluminum Chloride and Bromine gas. Single Replacement 3. Sodium Phosphate and Calcium Chloride react to form Calcium Phosphate and Sodium Chloride.

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## **Word Equations Answer Key Zinc and Lead (II) nitrate react ...**

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Oxidation-reduction reactions. Oxidation and reduction. Oxidation state trends in periodic table. Practice determining oxidation states. Unusual oxygen oxidation states. Oxidation-reduction (redox) reactions. This is the currently selected item. Introduction to redox reactions.

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Balancing Redox Reactions.

Oxidation/Reduction (Redox) reactions  
can be balanced using the oxidation



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state changes, as seen in the previous example. However, there is an easier method, which involves breaking a redox reaction into two half- reactions. This is best shown by working an example.

## **Worksheet 1 Redox Reactions - University Of Illinois**

Balancing redox reactions in acidic

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solution Problems #1-10 Ten Examples  
... That way leads to the correct answer  
without having to use half-reactions.  
There are some redox reactions where  
using half-reactions turns out to be  
"more" work, but there aren't that many.

**Balancing redox reactions in acidic  
solution: Problems #1-10**

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